Gas equilibria and Kp

AQA 7405 · Full A-level

Answer on your own paper. Show your working where needed.

1. Write a Kp expression

$PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$. Write the equilibrium expression for Kp using partial pressures. Deduce its units when every partial pressure is in kPa.

2. Go from amounts to partial pressures

$N_2O_4(g) \rightleftharpoons 2NO_2(g)$. At equilibrium there are 0.600 mol of N_2O_4 and 0.400 mol of NO_2. Total pressure is 150 kPa.

Find both partial pressures and Kp.

3. Two ways to write the Haber equilibrium

At equilibrium $p(N_2)$ = 75.0 kPa, $p(H_2)$ = 145 kPa and $p(NH_3)$ = 90.0 kPa for $N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$.

Find Kp and its units for the equation shown and for the equation written to form one mole of NH_3.

4. A vessel containing phosphorus chlorides

1.20 mol of $PCl_5(g)$ is present initially. At equilibrium for $PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$, 0.300 mol of PCl_5 has decomposed. Total pressure is 200 kPa.

Calculate Kp at this temperature.

5. Find a missing pressure

$CO(g) + H_2O(g) \rightleftharpoons CO_2(g) + H_2(g)$. Kp = 2.50. At equilibrium, p(CO) = 42.0 kPa, $p(H_2O)$ = 80.0 kPa and $p(H_2)$ = 70.0 kPa.

Find $p(CO_2)$.

6. Change pressure and temperature separately

$N_2O_4(g) \rightleftharpoons 2NO_2(g)$ is endothermic forwards. For each change, predict the equilibrium shift and what happens to Kp.

First consider compression at fixed temperature. Then consider heating the original mixture at fixed total pressure.

7. Approach the same equilibrium

Two vessels start with identical mixtures of $N_2O_4(g)$ and $NO_2(g)$ at the same pressure and temperature. Only one vessel contains a suitable catalyst.

What changes in the time taken to reach equilibrium, the equilibrium composition and Kp? Explain why.

8. Find the total equilibrium pressure

$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$. The equilibrium mole fractions of SO_2, O_2 and SO_3 are 0.250, 0.200 and 0.550 respectively. Kp is 0.242 kPa^{-1}.

Calculate the total pressure.

9. Extra challenge · Infer composition from total pressure

At fixed temperature and volume, a vessel initially contains 0.750 mol PCl_5 and 0.250 mol Ar at 125 kPa. For $PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$, the total pressure reaches 155 kPa at equilibrium. Argon does not react.

Use the change in total pressure to calculate Kp and give its units. Explain how argon enters the calculation.

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